

**RAKIRO BIOTECH SYSTEMS PVT LTD**

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Doc No : PDAE264**Date :** 01-02-2024**Type :** AQUASOL**Product Code:** AE264**PRODUCT DATA SHEET****1 INFORMATION**

CODE: AE264

PARAMETER: ACIDITY

RANGE: 10-200, 100-2000 mg/l as CaCO₃**2 METHOD**

Hydrogen ions present in a sample as a result of dissociation or hydrolysis of solutes react with additions of standard alkali. acidity thus depends on the end point ph or indicator used .

3 APPLICATION

Drinking Water, Mineral Water, Well Water, Swimming Pool Water, Surface and Ground Water, Aquaculture, Boiler Water, Process Water, Industrial Wastewater, Effluent Water, Cooling System Water, Chiller Water etc

4 INTERFERENCE

Dissolved gases contributing to acidity or alkalinity such as co₂ hydrogen sulfided, or ammonia, may be lost or gained during sampling, storage, or titration. Minimize such effect by titrating to the end point promptly after opening sample container, avoiding vigorous sheking or mixing , protecting sample from atmosphere during titration, and letting sample become nowarmer than it was at collection.do not use indicator with coloured or turbid sample that may obscure the colour change at the end point.

6 REAGENTS AND ACCESSORIES

Reagents: AK1(1Nos),AC2(1Nos), AC3(1Nos)

Accessories: 25ML Plastic Test Jar(1Nos), Plastic Spoon(1Nos), Procedure Label(1Nos)

7 STORAGE

The test reagents are stable up to the date stated on the pack when stored closed at ambient temperature

8 REFERENCE

APHA Standard Methods, 22nd ed., Method 2320-B– Standard Methods for Chemical Analysis of Water and Waste water. LARSON, T.E. & L.M. HENLEY 1955. Determination of low alkalinity or acidity in water. Anal Chem. 27:851.

9 DIRECTION FOR USE

1.Take 10 ml of water sample to be tested in the test jar.

2.Add 2 drops of AK 1 mix well to dissolve.

3.Now drop wise add AC 2 # counting the number of drops while mixing until the colour changes from colourless to pink.

* If the expected Acidity is more than 200 ppm, then use AC 3 instead of AC 2.

Calculations:

Total Acidity ppm as CaCO₃ = 10 X (Number of Drops of AC 2)

= 100 X (Number of Drops of AC 3)